

National curriculum or specification coverage

KS3 Chemistry: materials, states of matter, combustion, environmental chemistry

KS3 Chemistry: reactions, bonding, materials, combustion

5.1.1.1 Atoms, elements and compounds

5.1.1.2 Mixtures

5.1.1.3 The development of the model of the atom (common content with physics)

5.1.1.4 Relative electrical charges of subatomic particles

5.1.1.5 Size and mass of atoms

5.1.1.6 Relative atomic mass

5.1.1.7 Electronic structure

5.1.2.1 The periodic table

5.1.2.2 Development of the periodic table

5.1.2.3 Metals and non-metals

5.1.2.4 Group 0

5.1.2.5 Group 1

5.1.2.6 Group 7

5.2.1.1 Chemical bonds

5.2.1.2 Ionic bonding

5.2.1.3 Ionic compounds

5.2.1.4 Covalent bonding

5.2.1.5 Metallic bonding

Year 9 Long Term Scheme of Learning 2026-2027
Subject: Chemistry

Date	Exams/ Assessments	Learning Intentions	Core Task/Assessments
31/08/26		C8 Earth's Resources - Earth provides natural and artificial resources - Although water is renewable, only a tiny percentage is usable freshwater - Air is essential for breathing, combustion, and photosynthesis - Fossil fuels are non-renewable energy sources formed over millions of years - Materials such as ceramics, polymers, and composites have specific properties that make them useful, and recycling and sustainability help us manage resources responsibly for the future.	9C8 End of Unit Assessment
07/09/26			
14/09/26			
21/09/26			
28/09/26			
05/10/26			
12/10/26			
19/10/26			
26/10/26		HALF TERM	
02/11/26		C9 Reactivity & Electrolysis - Different metals react at different speeds - The reactivity series shows the order of metal reactivity - Metals below carbon in the reactivity series can be extracted by heating with carbon in a reduction reaction. - Metals above carbon must be extracted using electrolysis, which splits compounds using electricity.	9C9 End of Unit Assessment
09/11/26			
16/11/26			
23/11/26			
30/11/26			
07/12/26			
14/12/26			
21/12/26		CHRISTMAS HOLIDAY	
28/12/26			
04/01/27		- use the names and symbols of the first 20 elements in the periodic table, the elements in Groups 1 and 7, and other elements in this specification - name compounds of these elements from given formulae or symbol equations - write word equations for the reactions in this specification - write formulae and balanced chemical equations for the reactions in this specification. (HT only) write balanced half equations and ionic equations where appropriate. - describe, explain and give examples of the specified processes of separation - suggest suitable separation and purification techniques for mixtures when given appropriate information.	C1 standard task
11/01/27	Y9		
18/01/27	Y9		
25/01/27			
01/02/27			
08/02/27			
15/02/27		HALF TERM	
22/02/27		why the new evidence from the scattering experiment led to a change in the atomic model	C1 standard task – period table.
01/03/27			

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08/03/27		- the difference between the plum pudding model of the atom and the nuclear model of the atom. - How did Chadwick's work contribute to the understanding of isotopes. - Explain why atoms have no overall charge. - Explain what the atomic and mass number of the atoms are. - Use the nuclear model to describe atoms and isotopes. - calculate the numbers of protons, neutrons and electrons in an atom or ion, given its atomic number and mass number. - relate size and scale of atoms to objects in the physical world. - calculate the relative atomic mass of an element given the percentage abundance of its isotopes. - Explain how electronic structure of an atom can be represented.	
15/03/27			
22/03/27			
29/03/27	SPRING BREAK		
05/04/27			
12/04/27		- explain how the position of an element in the periodic table is related to the arrangement of electrons in its atoms and hence to its atomic number - predict possible reactions and probable reactivity of elements from their positions in the periodic table. - Describe how elements were arranged in early periodic tables. - Explain why Mendeleev's periodic table was accepted. - explain the differences between metals and non-metals on the basis of their characteristic physical and chemical properties. - explain how the atomic structure of metals and non-metals relates to their position in the periodic table - explain how the reactions of elements are related to the arrangement of electrons in their atoms and hence to their atomic number.	C1 standard task – extended response
19/04/27			
26/04/27			
03/05/27			
10/05/27			
17/05/27			
24/05/27			
31/05/27	HALF TERM		
07/06/27	Y9	- explain how properties of the elements in Group 0 depend on the outer shell of electrons of the atoms - predict properties from given trends down the group. - explain how properties of the elements in Group 7 depend on the outer shell of electrons of the atoms - predict properties from given trends down the group. - explain chemical bonding in terms of electrostatic forces and the transfer or sharing of electrons. - Describe and explain covalent, ionic and metallic bonding.	End of year assessment and feedback
14/06/27	Y9		
21/06/27			
28/06/27			
05/07/27			
12/07/27			