

Year 13 Long Term Scheme of Learning 2026-2027

Subject: Chemistry

National curriculum or specification coverage

AQA A-level Chemistry (7405) — Year 13 / A-level-only content

- 3.1.8 Thermodynamics
- 3.1.9 Rate equations
- 3.1.10 Equilibrium constant K_p for homogeneous systems
- 3.1.11 Electrode potentials and electrochemical cells
- 3.1.12 Acids and bases
- 3.2.4 Properties of Period 3 elements and their oxides
- 3.2.5 Transition metals
- 3.2.6 Reactions of ions in aqueous solution
- 3.3.7 Optical isomerism
- 3.3.8 Aldehydes and ketones
- 3.3.9 Carboxylic acids and derivatives
- 3.3.10 Aromatic chemistry
- 3.3.11 Amines
- 3.3.12 Polymers
- 3.3.13 Amino acids, proteins and DNA
- 3.3.14 Organic synthesis
- 3.3.15 Nuclear magnetic resonance spectroscopy
- 3.3.16 Chromatography
- Required practicals 7–12; relevant mathematical skills, apparatus and techniques, and practical endorsement competencies.

Date	Exams/ Assessments	Learning Intentions	Core Task/Assessments
31/08/2026		Thermodynamics; K_p ; rate equations; optical isomerism (3.1.8–3.1.10, 3.3.7) Construct Born–Haber cycles to determine lattice enthalpies and explain deviations from the perfect ionic model. Calculate enthalpy changes of solution using hydration and lattice enthalpies. Calculate entropy changes and Gibbs free-energy changes; interpret ΔG against temperature and determine feasibility. Calculate mole fractions, partial pressures and K_p , including units and effects of temperature. Determine initial and continuous rates, reaction orders, rate equations, rate constants and rate-determining steps; apply Arrhenius plots and calculations. Identify enantiomers and explain optical activity.	RP7: iodine-clock initial-rate method and continuous monitoring; rate graphs, uncertainties and Arrhenius calculations. K_p and Born–Haber problem sets; stereochemistry models. PPE1: 19/10/2026.

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Date	Exams/ Assessments	Learning Intentions	Core Task/Assessments
07/09/2026		Thermodynamics; K _p ; rate equations; optical isomerism (3.1.8–3.1.10, 3.3.7) Construct Born–Haber cycles to determine lattice enthalpies and explain deviations from the perfect ionic model. Calculate enthalpy changes of solution using hydration and lattice enthalpies. Calculate entropy changes and Gibbs free-energy changes; interpret ΔG against temperature and determine feasibility. Calculate mole fractions, partial pressures and K _p , including units and effects of temperature. Determine initial and continuous rates, reaction orders, rate equations, rate constants and rate-determining steps; apply Arrhenius plots and calculations. Identify enantiomers and explain optical activity.	RP7: iodine-clock initial-rate method and continuous monitoring; rate graphs, uncertainties and Arrhenius calculations. K _p and Born–Haber problem sets; stereochemistry models. PPE1: 19/10/2026.
14/09/2026		Thermodynamics; K _p ; rate equations; optical isomerism (3.1.8–3.1.10, 3.3.7) Construct Born–Haber cycles to determine lattice enthalpies and explain deviations from the perfect ionic model. Calculate enthalpy changes of solution using hydration and lattice enthalpies. Calculate entropy changes and Gibbs free-energy changes; interpret ΔG against temperature and determine feasibility. Calculate mole fractions, partial pressures and K _p , including units and effects of temperature. Determine initial and continuous rates, reaction orders, rate equations, rate constants and rate-determining steps; apply Arrhenius plots and calculations. Identify enantiomers and explain optical activity.	RP7: iodine-clock initial-rate method and continuous monitoring; rate graphs, uncertainties and Arrhenius calculations. K _p and Born–Haber problem sets; stereochemistry models. PPE1: 19/10/2026.
21/09/2026		Thermodynamics; K _p ; rate equations; optical isomerism (3.1.8–3.1.10, 3.3.7) Construct Born–Haber cycles to determine lattice enthalpies and explain deviations from the perfect ionic model. Calculate enthalpy changes of solution using hydration and lattice enthalpies. Calculate entropy changes and Gibbs free-energy changes; interpret ΔG against temperature and determine feasibility. Calculate mole fractions, partial pressures and K _p , including units and effects of temperature. Determine initial and continuous rates, reaction orders, rate equations, rate constants and rate-determining steps; apply Arrhenius plots and calculations. Identify enantiomers and explain optical activity.	RP7: iodine-clock initial-rate method and continuous monitoring; rate graphs, uncertainties and Arrhenius calculations. K _p and Born–Haber problem sets; stereochemistry models. PPE1: 19/10/2026.
28/09/2026		Thermodynamics; K _p ; rate equations; optical isomerism (3.1.8–3.1.10, 3.3.7) Construct Born–Haber cycles to determine lattice enthalpies and explain deviations from the perfect ionic model. Calculate enthalpy changes of solution using hydration and lattice enthalpies. Calculate entropy changes and Gibbs free-energy changes; interpret ΔG against temperature and determine feasibility. Calculate mole fractions, partial pressures and K _p , including units and effects of temperature. Determine initial and continuous rates, reaction orders, rate equations, rate constants and rate-determining steps; apply Arrhenius plots and calculations. Identify enantiomers and explain optical activity.	RP7: iodine-clock initial-rate method and continuous monitoring; rate graphs, uncertainties and Arrhenius calculations. K _p and Born–Haber problem sets; stereochemistry models. PPE1: 19/10/2026.
05/10/2026		Thermodynamics; K _p ; rate equations; optical isomerism (3.1.8–3.1.10, 3.3.7) Construct Born–Haber cycles to determine lattice enthalpies and explain deviations from the perfect ionic model. Calculate enthalpy changes of solution using hydration and lattice enthalpies. Calculate entropy changes and Gibbs free-energy changes; interpret ΔG against temperature and determine feasibility. Calculate mole fractions, partial pressures and K _p , including units and effects of temperature. Determine initial and continuous rates, reaction orders, rate equations, rate constants and rate-determining steps; apply Arrhenius plots and calculations. Identify enantiomers and explain optical activity.	RP7: iodine-clock initial-rate method and continuous monitoring; rate graphs, uncertainties and Arrhenius calculations. K _p and Born–Haber problem sets; stereochemistry models. PPE1: 19/10/2026.
12/10/2026		Thermodynamics; K _p ; rate equations; optical isomerism (3.1.8–3.1.10, 3.3.7) Construct Born–Haber cycles to determine lattice enthalpies and explain deviations from the perfect ionic model. Calculate enthalpy changes of solution using hydration and lattice enthalpies. Calculate entropy changes and Gibbs free-energy changes; interpret ΔG against temperature and determine feasibility. Calculate mole fractions, partial pressures and K _p , including units and effects of temperature. Determine initial and continuous rates, reaction orders, rate equations, rate constants and rate-determining steps; apply Arrhenius plots and calculations. Identify enantiomers and explain optical activity.	RP7: iodine-clock initial-rate method and continuous monitoring; rate graphs, uncertainties and Arrhenius calculations. K _p and Born–Haber problem sets; stereochemistry models. PPE1: 19/10/2026.

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19/10/2026	PPE1	Thermodynamics; K _p ; rate equations; optical isomerism (3.1.8–3.1.10, 3.3.7) Construct Born–Haber cycles to determine lattice enthalpies and explain deviations from the perfect ionic model. Calculate enthalpy changes of solution using hydration and lattice enthalpies. Calculate entropy changes and Gibbs free-energy changes; interpret ΔG against temperature and determine feasibility. Calculate mole fractions, partial pressures and K _p , including units and effects of temperature. Determine initial and continuous rates, reaction orders, rate equations, rate constants and rate-determining steps; apply Arrhenius plots and calculations. Identify enantiomers and explain optical activity.	RP7: iodine-clock initial-rate method and continuous monitoring; rate graphs, uncertainties and Arrhenius calculations. K _p and Born–Haber problem sets; stereochemistry models. PPE1: 19/10/2026.
02/11/2026		Transition metals and aqueous ions; carbonyls and carboxylic acid derivatives (3.2.5, 3.2.6, 3.3.8, 3.3.9) Explain variable oxidation states, complex-ion formation, coordination number, ligand substitution, shapes, colour and catalytic action of transition metals. Write relevant redox equations and calculate redox-titration results. Identify aqueous transition-metal ions from test-tube observations. Explain carbonyl functional groups and nucleophilic addition to aldehydes and ketones. Describe carboxylic acids, esters, acid chlorides and acid anhydrides, esterification, acylation and nucleophilic addition–elimination; connect fats and oils to esters.	RP11: identify transition-metal ions in aqueous solution. RP10a: prepare and test purity of an organic solid (aspirin); RP10b: prepare a pure organic liquid (ethyl ethanoate). Redox titration, synthesis and mechanism questions.
09/11/2026		Transition metals and aqueous ions; carbonyls and carboxylic acid derivatives (3.2.5, 3.2.6, 3.3.8, 3.3.9) Explain variable oxidation states, complex-ion formation, coordination number, ligand substitution, shapes, colour and catalytic action of transition metals. Write relevant redox equations and calculate redox-titration results. Identify aqueous transition-metal ions from test-tube observations. Explain carbonyl functional groups and nucleophilic addition to aldehydes and ketones. Describe carboxylic acids, esters, acid chlorides and acid anhydrides, esterification, acylation and nucleophilic addition–elimination; connect fats and oils to esters.	RP11: identify transition-metal ions in aqueous solution. RP10a: prepare and test purity of an organic solid (aspirin); RP10b: prepare a pure organic liquid (ethyl ethanoate). Redox titration, synthesis and mechanism questions.
16/11/2026		Transition metals and aqueous ions; carbonyls and carboxylic acid derivatives (3.2.5, 3.2.6, 3.3.8, 3.3.9) Explain variable oxidation states, complex-ion formation, coordination number, ligand substitution, shapes, colour and catalytic action of transition metals. Write relevant redox equations and calculate redox-titration results. Identify aqueous transition-metal ions from test-tube observations. Explain carbonyl functional groups and nucleophilic addition to aldehydes and ketones. Describe carboxylic acids, esters, acid chlorides and acid anhydrides, esterification, acylation and nucleophilic addition–elimination; connect fats and oils to esters.	RP11: identify transition-metal ions in aqueous solution. RP10a: prepare and test purity of an organic solid (aspirin); RP10b: prepare a pure organic liquid (ethyl ethanoate). Redox titration, synthesis and mechanism questions.
23/11/2026		Transition metals and aqueous ions; carbonyls and carboxylic acid derivatives (3.2.5, 3.2.6, 3.3.8, 3.3.9) Explain variable oxidation states, complex-ion formation, coordination number, ligand substitution, shapes, colour and catalytic action of transition metals. Write relevant redox equations and calculate redox-titration results. Identify aqueous transition-metal ions from test-tube observations. Explain carbonyl functional groups and nucleophilic addition to aldehydes and ketones. Describe carboxylic acids, esters, acid chlorides and acid anhydrides, esterification, acylation and nucleophilic addition–elimination; connect fats and oils to esters.	RP11: identify transition-metal ions in aqueous solution. RP10a: prepare and test purity of an organic solid (aspirin); RP10b: prepare a pure organic liquid (ethyl ethanoate). Redox titration, synthesis and mechanism questions.
30/11/2026		Transition metals and aqueous ions; carbonyls and carboxylic acid derivatives (3.2.5, 3.2.6, 3.3.8, 3.3.9) Explain variable oxidation states, complex-ion formation, coordination number, ligand substitution, shapes, colour and catalytic action of transition metals. Write relevant redox equations and calculate redox-titration results. Identify aqueous transition-metal ions from test-tube observations. Explain carbonyl functional groups and nucleophilic addition to aldehydes and ketones. Describe carboxylic acids, esters, acid chlorides and acid anhydrides, esterification, acylation and nucleophilic addition–elimination; connect fats and oils to esters.	RP11: identify transition-metal ions in aqueous solution. RP10a: prepare and test purity of an organic solid (aspirin); RP10b: prepare a pure organic liquid (ethyl ethanoate). Redox titration, synthesis and mechanism questions.

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07/12/2026		Transition metals and aqueous ions; carbonyls and carboxylic acid derivatives (3.2.5, 3.2.6, 3.3.8, 3.3.9) Explain variable oxidation states, complex-ion formation, coordination number, ligand substitution, shapes, colour and catalytic action of transition metals. Write relevant redox equations and calculate redox-titration results. Identify aqueous transition-metal ions from test-tube observations. Explain carbonyl functional groups and nucleophilic addition to aldehydes and ketones. Describe carboxylic acids, esters, acid chlorides and acid anhydrides, esterification, acylation and nucleophilic addition–elimination; connect fats and oils to esters.	RP11: identify transition-metal ions in aqueous solution. RP10a: prepare and test purity of an organic solid (aspirin); RP10b: prepare a pure organic liquid (ethyl ethanoate). Redox titration, synthesis and mechanism questions.
14/12/2026		Transition metals and aqueous ions; carbonyls and carboxylic acid derivatives (3.2.5, 3.2.6, 3.3.8, 3.3.9) Explain variable oxidation states, complex-ion formation, coordination number, ligand substitution, shapes, colour and catalytic action of transition metals. Write relevant redox equations and calculate redox-titration results. Identify aqueous transition-metal ions from test-tube observations. Explain carbonyl functional groups and nucleophilic addition to aldehydes and ketones. Describe carboxylic acids, esters, acid chlorides and acid anhydrides, esterification, acylation and nucleophilic addition–elimination; connect fats and oils to esters.	RP11: identify transition-metal ions in aqueous solution. RP10a: prepare and test purity of an organic solid (aspirin); RP10b: prepare a pure organic liquid (ethyl ethanoate). Redox titration, synthesis and mechanism questions.
04/01/2027		Transition metals and aqueous ions; carbonyls and carboxylic acid derivatives (3.2.5, 3.2.6, 3.3.8, 3.3.9) Explain variable oxidation states, complex-ion formation, coordination number, ligand substitution, shapes, colour and catalytic action of transition metals. Write relevant redox equations and calculate redox-titration results. Identify aqueous transition-metal ions from test-tube observations. Explain carbonyl functional groups and nucleophilic addition to aldehydes and ketones. Describe carboxylic acids, esters, acid chlorides and acid anhydrides, esterification, acylation and nucleophilic addition–elimination; connect fats and oils to esters.	RP11: identify transition-metal ions in aqueous solution. RP10a: prepare and test purity of an organic solid (aspirin); RP10b: prepare a pure organic liquid (ethyl ethanoate). Redox titration, synthesis and mechanism questions.
11/01/2027		Electrode potentials and cells; aromatic chemistry and amines (3.1.11, 3.3.10, 3.3.11) Write electrode half-equations and conventional cell representations; describe the standard hydrogen electrode and standard conditions. Calculate EMF, use electrode potentials to predict reaction direction and discuss limitations of feasibility predictions. Explain the chemistry of commercial electrochemical cells. Describe benzene bonding and stability, and compare electrophilic substitution with addition. Explain nitration and Friedel–Crafts reactions where specified. Name and classify amines; explain their basicity and preparation, including reduction of nitro compounds and nucleophilic substitution.	RP8: measure EMF of an electrochemical cell. Cell notation, electrode-potential and organic-mechanism questions. PPE2: 08/02/2027.
18/01/2027		Electrode potentials and cells; aromatic chemistry and amines (3.1.11, 3.3.10, 3.3.11) Write electrode half-equations and conventional cell representations; describe the standard hydrogen electrode and standard conditions. Calculate EMF, use electrode potentials to predict reaction direction and discuss limitations of feasibility predictions. Explain the chemistry of commercial electrochemical cells. Describe benzene bonding and stability, and compare electrophilic substitution with addition. Explain nitration and Friedel–Crafts reactions where specified. Name and classify amines; explain their basicity and preparation, including reduction of nitro compounds and nucleophilic substitution.	RP8: measure EMF of an electrochemical cell. Cell notation, electrode-potential and organic-mechanism questions. PPE2: 08/02/2027.
25/01/2027		Electrode potentials and cells; aromatic chemistry and amines (3.1.11, 3.3.10, 3.3.11) Write electrode half-equations and conventional cell representations; describe the standard hydrogen electrode and standard conditions. Calculate EMF, use electrode potentials to predict reaction direction and discuss limitations of feasibility predictions. Explain the chemistry of commercial electrochemical cells. Describe benzene bonding and stability, and compare electrophilic substitution with addition. Explain nitration and Friedel–Crafts reactions where specified. Name and classify amines; explain their basicity and preparation, including reduction of nitro compounds and nucleophilic substitution.	RP8: measure EMF of an electrochemical cell. Cell notation, electrode-potential and organic-mechanism questions. PPE2: 08/02/2027.

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Date	Exams/ Assessments	Learning Intentions	Core Task/Assessments
01/02/2027		Electrode potentials and cells; aromatic chemistry and amines (3.1.11, 3.3.10, 3.3.11) Write electrode half-equations and conventional cell representations; describe the standard hydrogen electrode and standard conditions. Calculate EMF, use electrode potentials to predict reaction direction and discuss limitations of feasibility predictions. Explain the chemistry of commercial electrochemical cells. Describe benzene bonding and stability, and compare electrophilic substitution with addition. Explain nitration and Friedel–Crafts reactions where specified. Name and classify amines; explain their basicity and preparation, including reduction of nitro compounds and nucleophilic substitution.	RP8: measure EMF of an electrochemical cell. Cell notation, electrode-potential and organic-mechanism questions. PPE2: 08/02/2027.
08/02/2027	PPE2	Electrode potentials and cells; aromatic chemistry and amines (3.1.11, 3.3.10, 3.3.11) Write electrode half-equations and conventional cell representations; describe the standard hydrogen electrode and standard conditions. Calculate EMF, use electrode potentials to predict reaction direction and discuss limitations of feasibility predictions. Explain the chemistry of commercial electrochemical cells. Describe benzene bonding and stability, and compare electrophilic substitution with addition. Explain nitration and Friedel–Crafts reactions where specified. Name and classify amines; explain their basicity and preparation, including reduction of nitro compounds and nucleophilic substitution.	RP8: measure EMF of an electrochemical cell. Cell notation, electrode-potential and organic-mechanism questions. PPE2: 08/02/2027.
22/02/2027		Electrode potentials and cells; aromatic chemistry and amines (3.1.11, 3.3.10, 3.3.11) Write electrode half-equations and conventional cell representations; describe the standard hydrogen electrode and standard conditions. Calculate EMF, use electrode potentials to predict reaction direction and discuss limitations of feasibility predictions. Explain the chemistry of commercial electrochemical cells. Describe benzene bonding and stability, and compare electrophilic substitution with addition. Explain nitration and Friedel–Crafts reactions where specified. Name and classify amines; explain their basicity and preparation, including reduction of nitro compounds and nucleophilic substitution.	RP8: measure EMF of an electrochemical cell. Cell notation, electrode-potential and organic-mechanism questions. PPE2: 08/02/2027.
01/03/2027		Electrode potentials and cells; aromatic chemistry and amines (3.1.11, 3.3.10, 3.3.11) Write electrode half-equations and conventional cell representations; describe the standard hydrogen electrode and standard conditions. Calculate EMF, use electrode potentials to predict reaction direction and discuss limitations of feasibility predictions. Explain the chemistry of commercial electrochemical cells. Describe benzene bonding and stability, and compare electrophilic substitution with addition. Explain nitration and Friedel–Crafts reactions where specified. Name and classify amines; explain their basicity and preparation, including reduction of nitro compounds and nucleophilic substitution.	RP8: measure EMF of an electrochemical cell. Cell notation, electrode-potential and organic-mechanism questions. PPE2: 08/02/2027.
08/03/2027		Period 3 and acids/bases; polymers, biological molecules and analysis (3.2.4, 3.1.12, 3.3.12–3.3.16) Explain Period 3 trends in melting point and bonding, reactions and acid–base behaviour of oxides. Apply Brønsted–Lowry acid–base theory; calculate pH, K_a , K_w , concentrations and buffer pH, including dilutions. Interpret titration curves and choose indicators. Explain addition and condensation polymerisation, hydrolysis and biodegradability. Describe amino acids, peptide bonds, proteins, enzymes and DNA. Use chromatography, NMR and chemical tests to determine structures and identify functional groups. Plan multistep organic syntheses using reaction pathways and appropriate conditions.	RP9: pH curves for weak acid/strong base and strong acid/weak base (full AQA scope). RP12: thin-layer chromatography. NMR and organic-synthesis practice; buffer calculations.
15/03/2027		Period 3 and acids/bases; polymers, biological molecules and analysis (3.2.4, 3.1.12, 3.3.12–3.3.16) Explain Period 3 trends in melting point and bonding, reactions and acid–base behaviour of oxides. Apply Brønsted–Lowry acid–base theory; calculate pH, K_a , K_w , concentrations and buffer pH, including dilutions. Interpret titration curves and choose indicators. Explain addition and condensation polymerisation, hydrolysis and biodegradability. Describe amino acids, peptide bonds, proteins, enzymes and DNA. Use chromatography, NMR and chemical tests to determine structures and identify functional groups. Plan multistep organic syntheses using reaction pathways and appropriate conditions.	RP9: pH curves for weak acid/strong base and strong acid/weak base (full AQA scope). RP12: thin-layer chromatography. NMR and organic-synthesis practice; buffer calculations.

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Date	Exams/ Assessments	Learning Intentions	Core Task/Assessments
22/03/2027		Period 3 and acids/bases; polymers, biological molecules and analysis (3.2.4, 3.1.12, 3.3.12–3.3.16) Explain Period 3 trends in melting point and bonding, reactions and acid–base behaviour of oxides. Apply Brønsted–Lowry acid–base theory; calculate pH, K_a , K_w , concentrations and buffer pH, including dilutions. Interpret titration curves and choose indicators. Explain addition and condensation polymerisation, hydrolysis and biodegradability. Describe amino acids, peptide bonds, proteins, enzymes and DNA. Use chromatography, NMR and chemical tests to determine structures and identify functional groups. Plan multistep organic syntheses using reaction pathways and appropriate conditions.	RP9: pH curves for weak acid/strong base and strong acid/weak base (full AQA scope). RP12: thin-layer chromatography. NMR and organic-synthesis practice; buffer calculations.
12/04/2027		Period 3 and acids/bases; polymers, biological molecules and analysis (3.2.4, 3.1.12, 3.3.12–3.3.16) Explain Period 3 trends in melting point and bonding, reactions and acid–base behaviour of oxides. Apply Brønsted–Lowry acid–base theory; calculate pH, K_a , K_w , concentrations and buffer pH, including dilutions. Interpret titration curves and choose indicators. Explain addition and condensation polymerisation, hydrolysis and biodegradability. Describe amino acids, peptide bonds, proteins, enzymes and DNA. Use chromatography, NMR and chemical tests to determine structures and identify functional groups. Plan multistep organic syntheses using reaction pathways and appropriate conditions.	RP9: pH curves for weak acid/strong base and strong acid/weak base (full AQA scope). RP12: thin-layer chromatography. NMR and organic-synthesis practice; buffer calculations.
19/04/2027		Period 3 and acids/bases; polymers, biological molecules and analysis (3.2.4, 3.1.12, 3.3.12–3.3.16) Explain Period 3 trends in melting point and bonding, reactions and acid–base behaviour of oxides. Apply Brønsted–Lowry acid–base theory; calculate pH, K_a , K_w , concentrations and buffer pH, including dilutions. Interpret titration curves and choose indicators. Explain addition and condensation polymerisation, hydrolysis and biodegradability. Describe amino acids, peptide bonds, proteins, enzymes and DNA. Use chromatography, NMR and chemical tests to determine structures and identify functional groups. Plan multistep organic syntheses using reaction pathways and appropriate conditions.	RP9: pH curves for weak acid/strong base and strong acid/weak base (full AQA scope). RP12: thin-layer chromatography. NMR and organic-synthesis practice; buffer calculations.
26/04/2027		Period 3 and acids/bases; polymers, biological molecules and analysis (3.2.4, 3.1.12, 3.3.12–3.3.16) Explain Period 3 trends in melting point and bonding, reactions and acid–base behaviour of oxides. Apply Brønsted–Lowry acid–base theory; calculate pH, K_a , K_w , concentrations and buffer pH, including dilutions. Interpret titration curves and choose indicators. Explain addition and condensation polymerisation, hydrolysis and biodegradability. Describe amino acids, peptide bonds, proteins, enzymes and DNA. Use chromatography, NMR and chemical tests to determine structures and identify functional groups. Plan multistep organic syntheses using reaction pathways and appropriate conditions.	RP9: pH curves for weak acid/strong base and strong acid/weak base (full AQA scope). RP12: thin-layer chromatography. NMR and organic-synthesis practice; buffer calculations.
03/05/2027		Revision and examination preparation (AQA 7405: all Year 12 and Year 13 content) Revisit physical, inorganic and organic chemistry across both years. Integrate synoptic calculations, mechanisms, spectroscopy, practical methods, data analysis, uncertainty and evaluation. Practise extended responses and timed Paper 1, Paper 2 and Paper 3 questions.	Past-paper questions, timed assessments, practical-skills review and targeted feedback.
10/05/2027		Revision and examination preparation (AQA 7405: all Year 12 and Year 13 content) Revisit physical, inorganic and organic chemistry across both years. Integrate synoptic calculations, mechanisms, spectroscopy, practical methods, data analysis, uncertainty and evaluation. Practise extended responses and timed Paper 1, Paper 2 and Paper 3 questions.	Past-paper questions, timed assessments, practical-skills review and targeted feedback.
17/05/2027		Revision and examination preparation (AQA 7405: all Year 12 and Year 13 content) Revisit physical, inorganic and organic chemistry across both years. Integrate synoptic calculations, mechanisms, spectroscopy, practical methods, data analysis, uncertainty and evaluation. Practise extended responses and timed Paper 1, Paper 2 and Paper 3 questions.	Past-paper questions, timed assessments, practical-skills review and targeted feedback.
24/05/2027		Revision and examination preparation (AQA 7405: all Year 12 and Year 13 content) Revisit physical, inorganic and organic chemistry across both years. Integrate synoptic calculations, mechanisms, spectroscopy, practical methods, data analysis, uncertainty and evaluation. Practise extended responses and timed Paper 1, Paper 2 and Paper 3 questions.	Past-paper questions, timed assessments, practical-skills review and targeted feedback.
07/06/2027		Revision and examination preparation (AQA 7405: all Year 12 and Year 13 content) Revisit physical, inorganic and organic chemistry across both years. Integrate synoptic calculations, mechanisms, spectroscopy, practical methods, data analysis, uncertainty and evaluation. Practise extended responses and timed Paper 1, Paper 2 and Paper 3 questions.	Past-paper questions, timed assessments, practical-skills review and targeted feedback.
14/06/2027		A-level examinations A-level / vocational examination period as indicated in the teaching plan. No additional new teaching content is specified.	Examinations and any scheduled school follow-up.

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Date	Exams/ Assessments	Learning Intentions	Core Task/Assessments
21/06/2027		A-level examinations A-level / vocational examination period as indicated in the teaching plan. No additional new teaching content is specified.	Examinations and any scheduled school follow-up.
28/06/2027		A-level examinations A-level / vocational examination period as indicated in the teaching plan. No additional new teaching content is specified.	Examinations and any scheduled school follow-up.